HONORS CHEMISTRY
HOMEWORK 11c. Name: _____

Find the pH of each of the following solutions:

1. 0.001 M HNO_3
2. 0.000001 M HCl
3. 0.1 M $HClO_3$
4. 0.1 M NaOH
5. 0.0001 M KOH
6. 1.0 M HBr

Use your calculator to find the pH of each of these solutions:

7. 0.45 M HCl
8. 2.33 M HNO_3
9. 0.00036 NaOH
10. 2.1 M KOH

Given the following conditions, find the $[H^+]$ or $[H_3O^+]$ concentration of each solution.

11. pH = 4.7
12. pH = 3.84
13. pH = 8.9
14. pOH = 6.4
15. pOH = 8.3
16. $[OH^-] = 5.2 * 10^{-4}$

17. Find the pOH of a solution whose hydronium ion concentration is $7.72 * 10^{-8}$

18. Find the $[OH^-]$ concentration of a solution whose pH is 8.115