

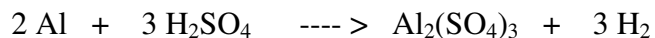
CHEMISTRY  
HOMEWORK 6f.

Name: \_\_\_\_\_

**A. REVIEW.** Write the correct formula of each compound:

- |                         |                         |                       |
|-------------------------|-------------------------|-----------------------|
| 1. potassium bromide    | 9. lithium sulfide      | 17. nitric acid       |
| 2. magnesium cyanide    | 10. sodium sulfate      | 18. nitrous acid      |
| 3. vanadium(III) oxide  | 11. copper(II) chlorate | 19. sulfuric acid     |
| 4. titanium(IV) oxide   | 12. calcium sulfite     | 20. sulfurous acid    |
| 5. ammonia              | 13. barium iodide       | 21. lithium phosphide |
| 6. ammonium fluoride    | 14. magnesium acetate   | 22. acetic acid       |
| 7. ammonium sulfate     | 15. hydrochloric acid   | 23. sulfur trioxide   |
| 8. nickel(II) phosphate | 16. rubidium carbonate  | 24. phosphoric acid   |

**B.** Aluminum reacts with sulfuric acid as follows:



If 250. g of aluminum react with a solution containing 600. g of  $\text{H}_2\text{SO}_4$ , find:

1. How many grams of aluminum sulfate will be formed.
2. How many moles of hydrogen gas will be formed.