CHEM. HOMEWORK 6c. Name: _____

A. REVIEW. Write the correct formula of each compound:

1. potassium bromide
2. magnesium cyanide
3. vanadium(III) oxide
4. titanium(IV) oxide
5. ammonia
6. ammonium fluoride
7. ammonium sulfate
8. nickel(II) phosphate
9. lithium sulfide
10. sodium sulfate
11. copper(II) chlorate
12. calcium sulfite
13. barium iodide
14. magnrsium acetate
15. hydrochloric acid
16. rubidium carbonate
17. nitric acid
18. nitrous acid
19. sulfuric acid
20. sulfurous acid
21. lithium phosphide
22. acetic acid
23. sulfur trioxide
24. phosphoric acid

B. Use the activity series shown on page 266 in your text, and the solubility rules, to determine whether or not a reaction will take place in each of the following cases.

If no reaction will occur (or only a balanced equilibrium), write NR or No Reaction.
If a reaction will occur, write a balanced equation for the reaction.

1.) $Cu + H_2SO_4$ ----- >

2.) $MgCl_2 + AgNO_3$ ----- >

3.) $NaBr + CuSO_4$ ----- >

4.) $Fe + Pb(NO_3)_2$ ----- >

5.) $Sn + KOH$ ----- >

6.) $Cl_2 + RbBr$ ----- >

7.) $(NH_4)_2SO_4 + Ba(C_2H_3O_2)_2$ ----- >

8.) $CaF_2 + Na_2SO_4$ ----- >

9.) $SnF_2 + Na$ ----- >

10.) $(NH_4)_2S + Al(ClO_3)_3$ ----- >