

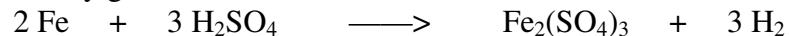
CHEMISTRY

HOMEWORK 6d.-> 5d. Name: _____ Sec. _____

1. Iron reacts with sulfuric acid, releasing hydrogen gas. If 80.0 g of iron are used, find:

a. how much hydrogen is released, and

b. how many grams of sulfuric acid are needed. The reaction is:



2. Aluminum reacts with a solution of zinc nitrate, producing aluminum nitrate and metallic zinc.

a. Write a balanced equation for this reaction.

b. What mass of zinc (in grams) is precipitated if 145 g of aluminum reacts?

3. Hot iron(III) oxide can react with carbon monoxide to form carbon dioxide and metallic iron.

a. write a balanced equation for this reaction

If 150. g of iron(III) oxide are used, find:

b. how many grams of iron can be obtained, and

c. how many moles of carbon monoxide are needed.